20. The molar conductivity of 0.025 M HCOOH (*aq*) soluti was measured as 46.1 S cm^2 mol^{-1}. Calculate its degr of dissociation and dissociation constant. [$\lambda°(H^+) = 349$ $\lambda°(HCOO^-) = 54.6$ S cm^2 mol^{-1}].

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21. Calculate $\Lambda°$ for $CaCl_2$ and $MgSO_4$ from the followi data : $\lambda°$ values are $Ca^{2+} = 119.0$, $Cl^- = 76.3$, $Mg^{2+} = 106$ $SO_4^{2-} = 160.0$ all in S cm^2 mol^{-1}.

***(NCERT Solved Example)**, (Pb. Board 202*

22. The conductivity of 0.001028 mol L^{-1} acetic acid is 4. $\times 10^{-5}$ S cm^{-1}. Calculate its dissociation constant if Λ (CH_3COOH) is 390.5 S cm^2 mol^{-1}.

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23. The molar conductances of NaOH, NaCl and $BaCl_2$ infinite dilution are 2.481×10^{-2}, 1.265×10^{-2} and 2.8 $\times 10^{-2}$ S m^2 mol^{-1} respectively. Calculate $\Lambda°_m$ for Ba(OH

24. The conductivity of a saturated solution of $BaSO_4$ 295 K, is found to be 3.758×10^{-6} ohm^{-1} cm^{-1} and th of water used is 1.36×10^{-6} ohm^{-1} cm^{-1}. Molar ion conductances at infinite dilution of Ba^{2+} and SO_4^{2-} ion are 110 and 136.6 ohm^{-1} cm^2 mol^{-1} respectively. Calcula the solubility of $BaSO_4$ at 295 K (Atomic masses: Ba = 13